



Cambridge IGCSE™ (9–1)

CO-ORDINATED SCIENCES

0973/11

Paper 1 Multiple Choice (Core)

October/November 2024

45 minutes

You must answer on the multiple choice answer sheet.

You will need: Multiple choice answer sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

INSTRUCTIONS

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A**, **B**, **C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

INFORMATION

- The total mark for this paper is 40.
- Each correct answer will score one mark.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.

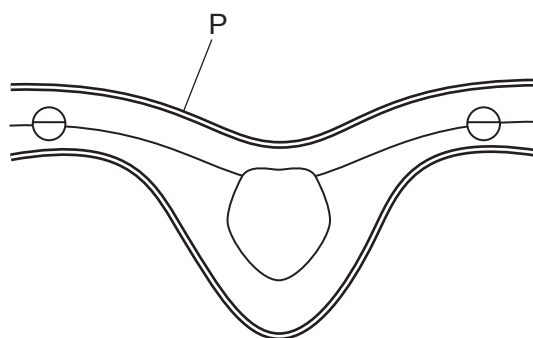
This document has **16** pages. Any blank pages are indicated.



- 4 One method of preventing food spoilage is to store it at 4 °C in a refrigerator.

Why does storing food at low temperatures help to prevent food spoilage?

- A It decreases enzyme activity.
 - B It denatures enzymes.
 - C It increases enzyme production.
 - D It kills cells.
- 5 The diagram shows a plan of the cross-section of a dicotyledonous leaf.



What is the name of the part labelled P?

- A palisade mesophyll
 - B spongy mesophyll
 - C upper epidermis
 - D vascular bundle
- 6 What is the dietary importance of fibre?
- A It builds and repairs tissue.
 - B It builds strong bones and teeth.
 - C It helps to prevent constipation.
 - D It provides a source of iron.
- 7 Selexipag is a drug used to treat a condition where the blood pressure going to the lungs is too high. The drug works by widening the blood vessel leading to the lungs to reduce the pressure.
- Which blood vessel going to the lungs is selexipag widening?
- A aorta
 - B pulmonary artery
 - C pulmonary vein
 - D vena cava

4

8 The list shows some processes that occur in the body.

- 1 cell division
- 2 oxygen diffusion
- 3 protein synthesis
- 4 temperature regulation

Which processes use energy released by respiration?

- A** 1, 3 and 4 **B** 1 and 2 **C** 2, 3 and 4 **D** 3 and 4 only

9 In which order does an impulse pass along neurones during a reflex action?

- A** motor → relay → sensory
- B** motor → sensory → relay
- C** sensory → motor → relay
- D** sensory → relay → motor

10 Before fertilisation in flowering plants, which structure must the pollen nucleus pass through to reach the nucleus in the ovule?

- A** filament
- B** petal
- C** stamen
- D** style

11 Albinism is a genetic condition which results in an absence of pigment in the skin, hair and eyes.

If two black mice produce an albino offspring, which Punnett square shows the correct cross?

A

	B	b
B	BB	Bb
B	BB	Bb

B

	B	b
B	BB	Bb
b	Bb	bb

C

	B	b
b	Bb	bb
b	Bb	bb

D

	b	b
b	bb	bb
b	bb	bb

key

B dominant allele for black

b recessive allele for albinism

12 What is the correct order for the food chain from the information given?

- beetles are herbivores
- mice eat the beetles
- owls are the tertiary consumers
- the leaves of a plant are the producer

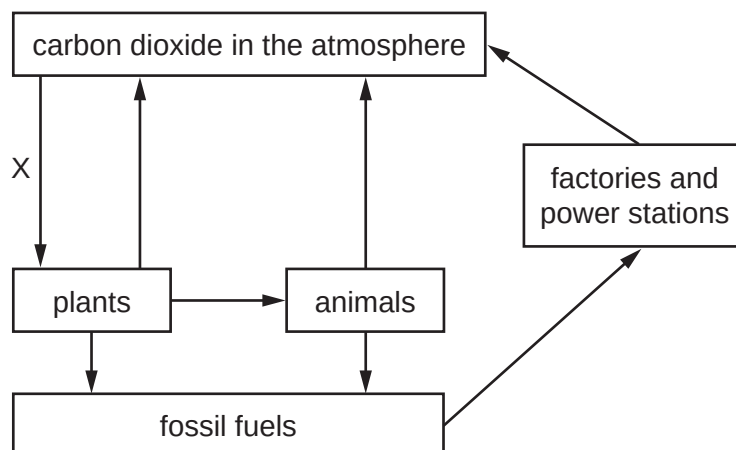
A beetles → mice → owl → leaves

B leaves → beetles → mice → owl

C leaves → beetles → owl → mice

D owl → mice → beetles → leaves

13 The diagram shows part of the carbon cycle.



Which process does X represent?

- A combustion
- B decomposition
- C photosynthesis
- D respiration

14 A student adds 5 g of magnesium ribbon to 20 cm³ of dilute hydrochloric acid in a beaker.

The student measures how long it takes for the effervescence to stop.

Which pieces of apparatus does the student use in this experiment?

	balance	measuring cylinder	stop-clock	thermometer
A	✓	✓	✗	✓
B	✓	✗	✓	✓
C	✓	✓	✓	✗
D	✗	✓	✓	✓

key

✓ = uses the apparatus

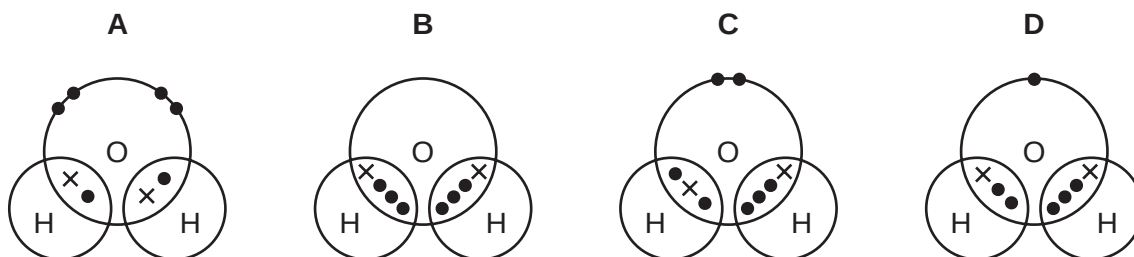
✗ = does not use the apparatus

15 Sea water is a mixture that contains sodium chloride dissolved in water.

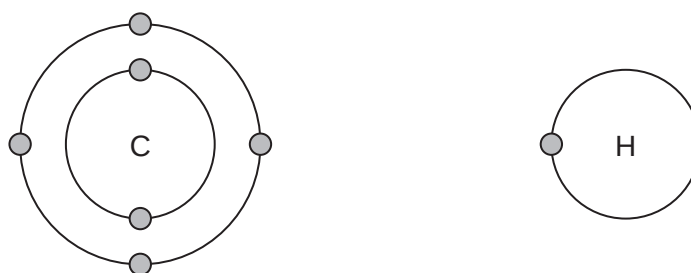
Which row describes sea water?

	solvent	solute	solution
A	water	sea water	sodium chloride
B	sea water	sodium chloride	water
C	water	sodium chloride	sea water
D	sodium chloride	water	sea water

16 What is the dot-and-cross diagram for a water molecule?



17 The electronic structures of carbon and of hydrogen are shown.



What is the formula of a compound formed between carbon and hydrogen?

- A** CH₂ **B** CH₃ **C** C₄H **D** CH₄

18 Which elements are formed at the electrodes during the electrolysis of concentrated aqueous sodium chloride?

	anode	cathode
A	chlorine	sodium
B	chlorine	hydrogen
C	hydrogen	chlorine
D	sodium	chlorine

- 19** Four different substances are added to the same acid of the same concentration in reactions W, X, Y and Z.

The initial temperature of the acid before each reaction is 21 °C.

The final temperatures of the mixtures are measured.

The results are shown.

reaction	W	X	Y	Z
final temperature / °C	28	19	26	17

Which row is correct?

	most endothermic reaction	most exothermic reaction
A	W	Z
B	Z	W
C	X	Y
D	Y	X

- 20** Dilute hydrochloric acid reacts with solid calcium carbonate.

Which change decreases the rate of the reaction?

- A** decreasing the concentration of the hydrochloric acid
- B** decreasing the size of the calcium carbonate pieces
- C** increasing the surface area of the calcium carbonate
- D** increasing the temperature of the acid

- 21** Which statement about oxidation is correct?

- A** It occurs when an element or compound chemically combines with oxygen.
- B** It occurs when an element or compound forms a mixture with oxygen.
- C** It occurs when an element or compound is separated from a mixture with oxygen.
- D** It occurs when a compound loses oxygen.

22 Which substance is classified as an acidic oxide?

- A calcium oxide
- B carbon dioxide
- C lithium oxide
- D sodium oxide

23 The results of two tests on compound P are shown.

test	observation
flame test	red flame
addition of acidified silver nitrate	white precipitate

What is P?

- A lithium bromide
- B lithium chloride
- C sodium bromide
- D sodium chloride

24 What is **not** a property of transition elements?

- A form coloured compounds
- B good electrical conductivity
- C high melting point
- D low density

25 What is the order of reactivity of metals, from highest to lowest?

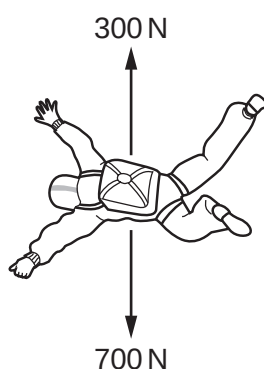
- A aluminium → calcium → copper → iron
- B calcium → iron → copper → aluminium
- C calcium → magnesium → iron → copper
- D aluminium → magnesium → iron → copper

- 26** A colourless liquid is tested with blue cobalt(II) chloride paper.

The blue paper turns pink.

What does this test prove that the liquid contains?

- A** an acid
 - B** an alkali
 - C** ethanol
 - D** water
- 27** What are the products of the complete combustion of ethanol?
- A** carbon and hydrogen
 - B** carbon dioxide and hydrogen
 - C** carbon dioxide and water
 - D** carbon monoxide and water
- 28** A person runs 5000 m in 1200 s and then walks a further 3000 m in 1800 s.
- What is their average speed for this journey?
- A** 1.7 m/s **B** 2.7 m/s **C** 2.9 m/s **D** 5.8 m/s
- 29** The diagram shows the two forces acting on a skydiver.



What is the resultant force on the skydiver?

- A** 400 N downwards
- B** 400 N upwards
- C** 1000 N downwards
- D** 1000 N upwards

- 30** When a ball is kicked vertically upwards, its initial kinetic energy is 12 J. When the ball arrives back at its starting point, its kinetic energy is 8.0 J.

Which statement explains the change in the kinetic energy of the ball?

- A** 4.0 J of work is done to increase the gravitational potential energy of the ball.
- B** 4.0 J of work is done to overcome air resistance.
- C** 20 J of work is done to increase the gravitational potential energy of the ball.
- D** 20 J of work is done to overcome air resistance.

- 31** A motor lifts a load vertically upwards at constant speed.

The weight of the load is known.

Which of the other quantities in the table must be known to determine the useful power output of the motor?

	height lifted	time to lift the load
A	✓	✓
B	✓	x
C	x	✓
D	x	x

key

✓ = this quantity is needed

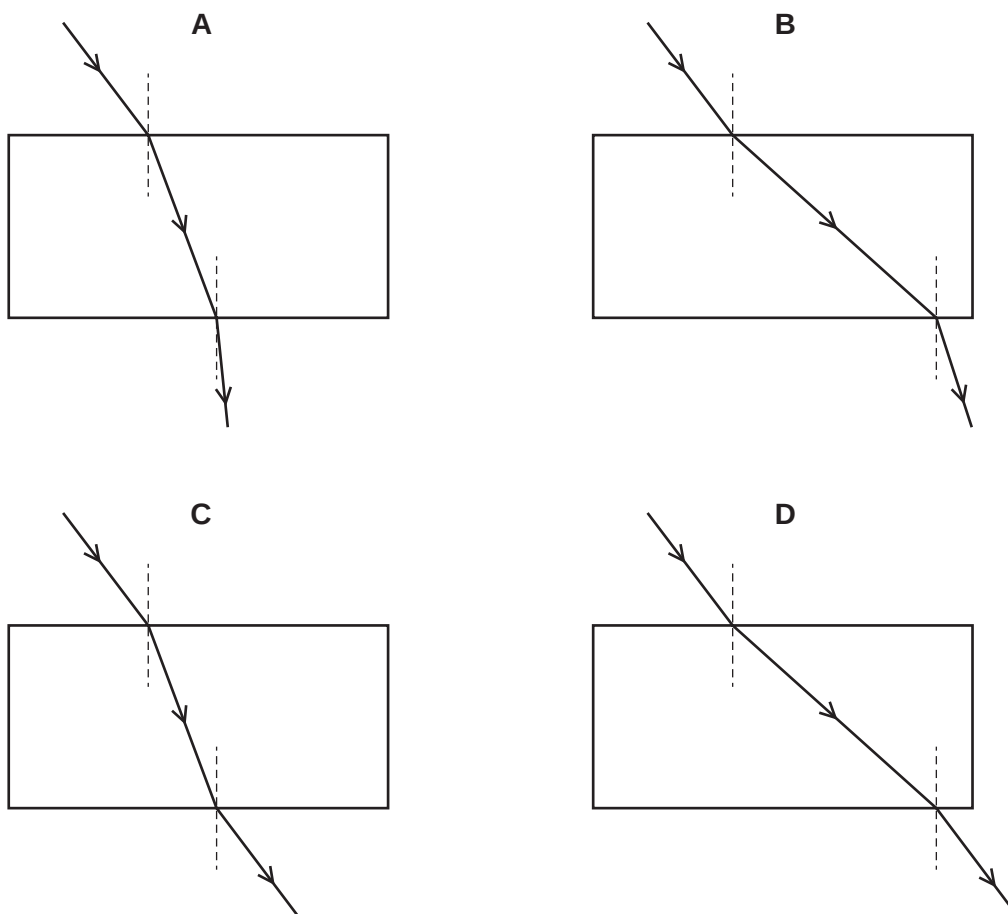
x = this quantity is **not** needed

- 32** Which material is a poor thermal conductor?

- A** air
- B** aluminium
- C** copper
- D** iron

33 A ray of light travels from air into a parallel-sided glass block and then back into air.

Which diagram shows the path of the light?



34 Which electromagnetic waves have the lowest frequency?

- A** gamma
- B** infrared
- C** ultraviolet
- D** radio

35 Object P is positively charged and object Q is negatively charged.

Which row describes the force between P and Q and explains how they became charged?

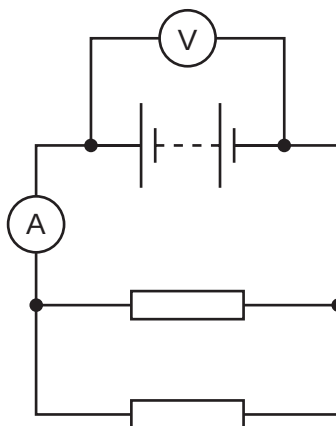
	force between P and Q	charge on P caused by	charge on Q caused by
A	attractive	gaining electrons	losing electrons
B	attractive	losing electrons	gaining electrons
C	repulsive	gaining electrons	losing electrons
D	repulsive	losing electrons	gaining electrons

- 36** A lamp is connected to a 3.0 V battery. The resistance of the lamp is $60\ \Omega$.

What is the current in the lamp?

- A** 0.050 mA **B** 20 mA **C** 50 mA **D** 180 mA

- 37** The diagram shows two resistors connected in parallel to a battery. The circuit contains a voltmeter and an ammeter.

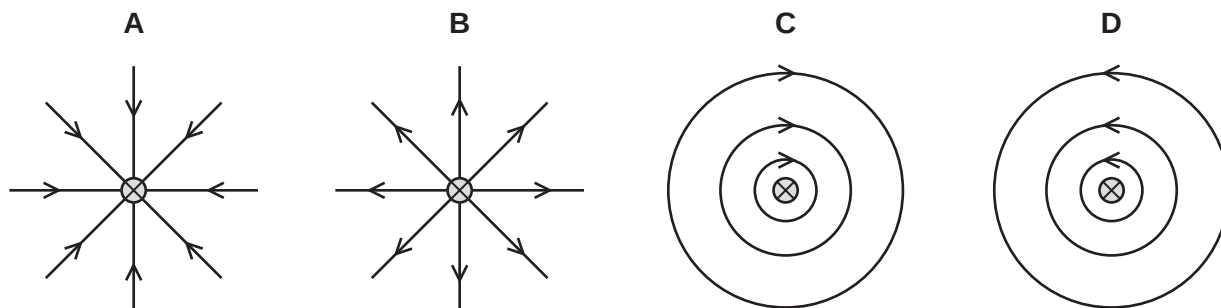


A third resistor is now connected in parallel with these two resistors.

What happens to the voltmeter reading and what happens to the ammeter reading?

	ammeter reading	voltmeter reading
A	decreases	increases
B	decreases	stays the same
C	increases	increases
D	increases	stays the same

- 38** Which diagram shows the pattern and the direction of the magnetic field around a straight wire that is carrying a current into the page?

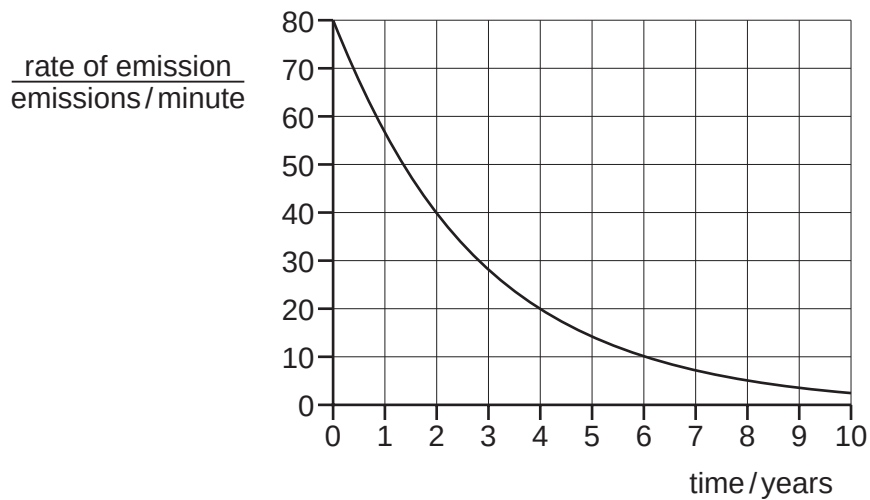


- 39 An element has two isotopes.

Which statement about the nuclei of atoms of these isotopes is correct?

- A They have equal numbers of neutrons but different numbers of electrons.
- B They have equal numbers of neutrons but different numbers of protons.
- C They have equal numbers of protons but different numbers of electrons.
- D They have equal numbers of protons but different numbers of neutrons.

- 40 The graph shows how the rate of emission of radiation from a radioactive sample changes with time.



What is the half-life of this sample?

- A 40 minutes
- B 2.0 years
- C 5.0 years
- D 10 years

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The Periodic Table of Elements

Group																			
I	II											III	IV	V	VI	VII	VIII		
		1 H hydrogen 1																	
		Key atomic number atomic symbol name relative atomic mass																	
3 Li lithium 7	4 Be beryllium 9													5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	10 Ne neon 20
11 Na sodium 23	12 Mg magnesium 24													13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40
19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84		
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium —	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131		
55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids		72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	75 Re rhenium 186	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —	85 At astatine —	86 Rn radon —	
87 Fr francium —	88 Ra radium —	89–103 actinoids		104 Rf rutherfordium —	105 Db dubnium —	106 Sg seaborgium —	107 Bh bohrium —	108 Hs hassium —	109 Mt meitnerium —	110 Ds darmstadtium —	111 Rg roentgenium —	112 Cn copernicium —	113 Nh nihonium —	114 Fl flerovium —	115 Mc moscovium —	116 Lv livermorium —	117 Ts tennessine —	118 Og oganesson —	

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
	89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —

The volume of one mole of any gas is 24dm³ at room temperature and pressure (r.t.p.).